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14. Use the Gas Law
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L of gas with 2 moles

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of helium gas and 3 moles of neon gas when the temperature is 300K? 15. Use the Gas Law Simulation to find how the pressure in the last example changes when the total volume is increased to 20.0L?

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Boyle's Law, simply

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stated: If the amount of gas and the temperature of the gas are held constant, the PRESSURE is inversely related or proportional to the VOLUME. ...
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14. Use the Gas Law Simulation to find the pressure (lock the pressure) when there is a total volume of 10.0

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f. What is the formula of the law? g. Draw a before and after picture of the gas law experiment showing what was manipulated and what was the effect (make sure you show gas particles) h.

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Give a real-life example of the law and briefly explain how it works. 5. Complete “Behavior of Gases” worksheet # 1, 3, 5 6.

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atm = 760.0 mm Hg =
101.3 kPa = 760 .0 torr

Boyle's Law Problems:

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?
2. A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.

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Molecular Theory will help you to explain things like how a hot air balloon rises and why car tires get a little flat in the winter.

WebQuest: Kinetic Molecular Theory and Gas Laws

2 Unit 2 Packet: Gas Laws Introduction to Gas Laws Notes: In chemistry, the relationships between gas physical properties are described as gas

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laws. Some of these properties are pressure, volume, and temperature. These laws show how a change in one of these properties affects the others.

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